Acta Crystallographica Section E

Structure Reports Online

ISSN 1600-5368

Sema Öztürk Yıldırım,^a Mehmet Akkurt,^a* Ülkü Yılmaz,^b Hasan Kücükbay^b and Vickie McKee^c

^aDepartment of Physics, Faculty of Arts and Sciences, Erciyes University, 38039 Kayseri, Turkey, ^bDepartment of Chemistry, Faculty of Arts and Sciences, Ínönü University, 44280 Malatya, Turkey, and ^cDepartment of Chemistry, Loughborough University, Leicestershire LE11 3TU, England

Correspondence e-mail: akkurt@erciyes.edu.tr

Key indicators

Single-crystal X-ray study T = 150 KMean $\sigma(\text{C-C}) = 0.003 \text{ Å}$ R factor = 0.033 wR factor = 0.089Data-to-parameter ratio = 10.2

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

5,5a-Diallyl-5,5a,13,14-tetrahydro-12*H*-di-1,3-benzimidazolo[1,2-*a*;1',2'-*c*][1,4]diazepine

The title compound, $C_{23}H_{24}N_4$, was synthesized from 3,3'-diallyl-1,1'-propylenedi(benzimidazole) dibromide and NaH in tetrahydrofuran solution. In the molecule, the diazepine ring exhibits a boat conformation.

Received 10 November 2006 Accepted 14 November 2006

Comment

Electron-rich olefins have attracted considerable attention in both the organic and inorganic preparative literature as a result of their unique properties as reagents and reaction intermediates (Böhm & Herrmann, 2000). They have been used as powerful reducing agents (Lappert, 1988), sources of carbene transition metal complexes (Küçükbay et al., 1996) and catalysts for acyloin type C-C coupling reactions (Çetinkaya & Küçükbay, 1995). They have an extensive chemistry and, in particular, electron-rich olefins that contain an imidazolidine or benzothiazolidine group have long been known, although there are few studies of electron-rich olefins containing a benzimidazolidine group. Isolation of allyl-, crotyl- or benzyl-substituted electron-rich olefins tends to be difficult because the synthesized olefins spontaneously transform to their [1,3]-sigmatropic rearrangement products. As was previously reported (Baldwin & Walker, 1974; Baldwin et al., 1977; Cetinkaya et al., 1998), we also obtained a [1,3]sigmatropic rearrangement product, namely 2',3'-diallyl-2',3'H-dibenzimidazolo[a,c]perhydro-1,4-diazepine, instead of the corresponding electron-rich olefin, (1), from a 3,3'-bis(allyl)-1,1'-propylendi(benzimidazole) dibromide and NaH in THF solution. The crystal structure of (2) is presented here.

The molecular structure of (2) is shown in Fig. 1. The geometric parameters in (2) are within the normal ranges (Allen *et al.*, 1987) and agree with those in similar structures reported in the literature (Mague & Eduok, 2000; Akkurt *et al.*, 2006*a,b*). The diazepine ring exhibits a boat conformation. The displacements of atoms N3, C17 and C8 from the C1/N2/C9/C10 mean plane are 0.398 (1), 0.214 (2) and 0.562 (2) Å, respectively. The benzimidazole ring systems in (2) are essentially planar and the dihedral angle between them is 75.56 (5)°.

© 2006 International Union of Crystallography All rights reserved

organic papers

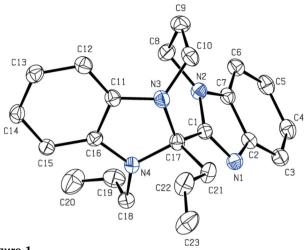


Figure 1

Molecular structure of (2), with the atom-numbering scheme. Displacement ellipsoids are drawn at the 50% probability level. For clarity, H atoms have been omitted.

The molecular conformation of (2) is stabilized by an intramolecular $C-H\cdots N$ hydrogen-bonding interaction (Table 2).

Experimental

A mixture of 3,3'-bis(allyl)-1,1'-propylendi(benzimidazole) dibromide (5.0 g, 9.7 mmol) and NaH (0.5 g, 21 mmol) in THF (50 ml) was stirred for 10 h at room temperature. Volatiles were eliminated *in vacuo*, toluene (20 ml) was added and the suspension was filtered. The resulting bright-yellow filtrate was concentrated to *ca* 10 ml and *n*-hexane (10 ml) was added. Upon cooling, colourless crystals of (2) (2 g, 53%) were obtained (m.p. 408–409 K). ¹H NMR (CDCl₃): δ 1.6–1.8 (m, -CH₂-bridge, 2H), 3.0–3.2 (d, -CH₂-, 2H), 3.7–4.0 (m, -N–CH₂-bridge, 4H), 4.2 (d, N–CH₂-, 2H), 4.8–5.0 (q, =-CH₂, 2H), 5.1–5.3 (q, =-CH₂, 2H), 5.4–5.6 (m, -CH=, 1H), 5.9–6.1 (m, -CH=, 1H), 6.2–7.8 (m, Ar–H, 8H). Analysis calculated for C₂₃H₂₄N₄: C 77.53, H 6.74, N 15.73%; found: C 76.62, H 6.98, N 16.40%.

Crystal data

Data collection

 $\begin{array}{lll} \text{Bruker SMART CCD area-detector} & 182 \\ \text{diffractometer} & 249 \\ \varphi \text{ and } \omega \text{ scans} & 233 \\ \text{Absorption correction: multi-scan} & R_{\text{in}} \\ (SADABS; \text{Sheldrick, 2003}) & \theta_{\text{max}} \\ T_{\text{min}} = 0.967, T_{\text{max}} = 0.970 \\ \end{array}$

18216 measured reflections 2499 independent reflections 2334 reflections with $I > 2\sigma(I)$ $R_{\rm int} = 0.030$ $\theta_{\rm max} = 28.6^{\circ}$

Refinement

 $\begin{array}{lll} \text{Refinement on } F^2 & w = 1/[\sigma^2(F_{\text{o}}^2) + (0.0508P)^2 \\ R[F^2 > 2\sigma(F^2)] = 0.033 & + 0.4382P] \\ wR(F^2) = 0.089 & \text{where } P = (F_{\text{o}}^2 + 2F_{\text{c}}^2)/3 \\ S = 1.02 & (\Delta/\sigma)_{\text{max}} < 0.001 \\ 2499 \text{ reflections} & \Delta\rho_{\text{max}} = 0.48 \text{ e Å}^{-3} \\ 244 \text{ parameters} & \Delta\rho_{\text{min}} = -0.19 \text{ e Å}^{-3} \\ \text{H-atom parameters constrained} \end{array}$

Table 1 Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	D-H	$H \cdot \cdot \cdot A$	$D \cdot \cdot \cdot A$	$D-H\cdots A$
C22—H22···N3	0.93	2.56	2.912 (4)	103

H atoms were placed in geometrically idealized positions and constrained to ride on their parents atoms, with C—H = 0.93–0.97 Å, and with $U_{\rm iso}({\rm H})=1.2U_{\rm eq}({\rm C})$. In the absence of significant anomalous scattering effects, Friedel pairs were merged.

Data collection: *APEXII* (Bruker, 2003); cell refinement: *SAINT* (Bruker, 1998); data reduction: *SAINT*; program(s) used to solve structure: *SIR97* (Altomare *et al.*, 1999); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

ÜY and HK thank İnönü University Research Fund (BAPB-2006- 41 and directed project BAPB-2006-11) for financial support for this study.

References

Akkurt, M., Yıldırım, S. Ö., Küçükbay, H., Şireci, N. & Büyükgüngör, O. (2006a). Acta Cryst. E62, o3184–o3186.

Akkurt, M., Yıldırım, S. Ö., Küçükbay, H., Şireci, N. & Büyükgüngör, O. (2006b). Acta Cryst. E62, o3512–o3514.

Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). J. Chem. Soc. Perkin Trans. 2, pp. S1–19.

Altomare, A., Burla, M. C., Camalli, M., Cascarano, G. L., Giacovazzo, C., Guagliardi, A., Moliterni, A. G. G., Polidori, G. & Spagna, R. (1999). J. Appl. Cryst. 32, 115–119.

Baldwin, J. E., Branz, S. E. & Walker, J. A. (1977). J. Org. Chem. 42, 4142–4144. Baldwin, J. E. & Walker, J. A. (1974). J. Am. Chem. Soc. 23, 596–597.

Böhm, V. P. W. & Herrmann, W. A. (2000). Angew. Chem. Int. Ed. 39, 4036–4038.

Bruker (1998). SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.

Bruker (2003). APEXII. Bruker AXS Inc., Madison, Wisconsin, USA.

Çetinkaya, B., Çetinkaya, E., Chamizo, J. A., Hitchcock, P. B., Jasim, H. A., Küçükbay, H. & Lappert, M. F. (1998). *J. Chem. Soc. Perkin Trans.* 1, pp. 2047–2054.

Çetinkaya, E. & Küçükbay, H. (1995). Turk. J. Chem. 19, 24-30.

Farrugia, L. J. (1997). J. Appl. Cryst. 30, 565.

Farrugia, L. J. (1999). J. Appl. Cryst. 32, 837-838.

Küçükbay, H., Çetinkaya, B., Guesmi, S. & Dixneuf, P. H. (1996). Organometallics, 15, 2434–2439.

Lappert, M. F. (1988). J. Organomet. Chem. 358, 185-214.

Mague, J. T. & Eduok, E. E. (2000). J. Chem. Crystallogr. 30, 311-313.

Sheldrick, G. M. (1997). SHELXL97. University of Göttingen, Germany.

Sheldrick, G. M. (2003). SADABS. Version 2.10. Bruker AXS Inc., Madison, Wisconsin, USA.